

## HYDROGEN

- Q.1** Hydrogen is evolved by the action of cold dil.  $\text{HNO}_3$  on-
- (A) Fe                      (B) Mn                      (C) Cu                      (D) Al
- Q.2** HCl is added to following oxides. Which one would give  $\text{H}_2\text{O}_2$  ?
- (A)  $\text{MnO}_2$                       (B)  $\text{PbO}_2$   
(C) BaO                      (D) None of these
- Q.3** Which one of the following chlorides will not fume in air ?
- (A)  $\text{BiCl}_3$                       (B)  $\text{CCl}_4$   
(C)  $\text{PCl}_5$                       (D) None of the air
- Q.4** Hydrogen will not reduce-
- (A) Heated cupric oxide                      (B) Heated ferric oxide  
(C) Heated stannic oxide                      (D) Heated aluminium oxide
- Q.5** The oxide that gives hydrogen peroxide on treatment with dilute acid is-
- (A)  $\text{PbO}_2$                       (B)  $\text{Na}_2\text{O}_2$                       (C)  $\text{MnO}_2$                       (D)  $\text{TiO}_2$
- Q.6** The sum of the number of neutrons and proton in the isotope of hydrogen is-
- (A) 6                      (B) 5                      (C) 4                      (D) 3
- Q.7** Which of the following pair will not produce dihydrogen gas ?
- (A)  $\text{Fe} + \text{H}_2\text{SO}_4$                       (B)  $\text{Mg} + \text{Steam}$   
(C)  $\text{Na} + \text{alcohol}$                       (D)  $\text{Cu} + \text{HCl (dil.)}$
- Q.8** The amount of  $\text{H}_2\text{O}_2$  present in 1L of 1.5 N  $\text{H}_2\text{O}_2$  solution is-
- (A) 2.5 g                      (B) 25.5 g                      (C) 3.0 g                      (D) 8.0 g

**Q.9** The maximum possible number of hydrogen bonds a water molecule can form is-

(A) 1

(B) 2

(B) 3

(D) 4

**Q.10** Hydrolysis of one mole of peroxodisulphuric acid produces-

(A) Two moles of sulphuric acid

(B) Two moles of peroxomonosulphuric acid

(C) One mole of sulphuric acid and one mole of peroxomonosulphuric acid

(D) One mole of sulphuric acid, one mole of peroxomonosulphuric acid and one mole of hydrogen peroxide.