

Dear student following is Tough level [●●●○○] test paper. Any 15 correct attempts in 25 minutes would be a satisfactory performance. Many questions have narrow difference among choices be careful.

SINGLE OPTION CORRECT :

- Q.1** During qualitative test of nitrate radical, a brown ring is formed. The ring formed is due to the formation of :
 (A) FeSO_4NO (B) $(\text{FeSO}_4)_2\text{NO}$
 (C) $\text{FeSO}_4(\text{NO})_2$ (D) $[\text{Fe}(\text{H}_2\text{O})_5\text{NO}] \text{SO}_4$
- Q.2** Evolution of deep red vapours, when an inorganic salt is mixed with powdered $\text{K}_2\text{Cr}_2\text{O}_7$ and heated with conc. H_2SO_4 , indicates the presence of :
 (A) Cl^- (B) F^-
 (C) BO_3^{3-} (D) PO_4^{3-}
- Q.3** Which of the following salts is/are not decomposed by dilute or conc. H_2SO_4 ?
 (A) Na_2SO_3 (B) NaCl
 (C) Na_2SO_4 (D) $(\text{CH}_3\text{COO})_2\text{Ca}$
- Q.4** Which of the following salts will turn water coloured when fumes evolved on treatment with conc. H_2SO_4 are passed in water ?
 (A) Nitrate (B) Bromide
 (C) Both (D) None
- Q.5** Starch iodide paper is used to test the presence of :
 (A) Iodine (B) Iodide ion
 (C) Oxidising agent (D) Reducing agent
- Q.6** In which of the following tests, Cl^- ion is tested by observing the colour of a precipitate which does not contain Cl^- ions at all ?
 (A) Chromyl chloride test
 (B) Lassaigne's test
 (C) Silver mirror test
 (D) Fehling solution test
- Q.7** For the test of halides, the sodium carbonate extract is acidified by :
 (A) dil. H_2SO_4 (B) dil. HNO_3
 (C) dil. HCl (D) Any of these
- Q.8** An inorganic mixture, when treated with dil. H_2SO_4 produced a gas which turned lead acetate paper black and potassium dichromate paper green. The mixture contains :
 (A) S^{2-} ions (B) SO_3^{2-} ions
 (C) NO_2^- ions (D) CH_3COO^- ions
- Q.9** When an iodide salt is treated with conc. H_2SO_4 , the product formed is :
 (A) I_2 (B) HI
 (C) IO_3^- (D) IO^-
- Q.10** An aqueous solution containing acetate ions, gives blood red colour with neutral FeCl_3 . On boiling, a ppt. is formed. The ppt. is :
 (A) Ferric acetate
 (B) Basic ferric acetate
 (C) Ferric oxide
 (D) Ferrous acetate
- Q.11** An inorganic mixture is heated with conc. H_2SO_4 . The evolved gas burns with blue flame. This is due to the presence of :
 (A) $\text{C}_2\text{O}_4^{2-}$ (B) Ba^{2+}
 (C) Cu^{2+} (D) BO_3^{3-}
- Q.12** Sodium carbonate extract of an inorganic salt is acidified with CH_3COOH and then added to CaCl_2 solution. The ppt. obtained decolourised acidified KMnO_4 . Probable inorganic salt is :
 (A) A carbonate (B) An oxalate
 (C) A sulphate (D) A sulphite
- Q.13** An inorganic salt, when treated with conc. H_2SO_4 , produced oily drops. The probable salt is :
 (A) An oxalate (B) A fluoride
 (C) A nitrite (D) An iodide



- Q.14** An inorganic mixture was treated with excess of dil. H_2SO_4 . Effervescences were produced. The solution was heated till effervescences ceased. After this a small pinch of MnO_2 were added. Fresh effervescences were produced. Select the correct statement/s :
- (A) Mixture contains CO_3^{2-} ions
 (B) Mixture contains $\text{C}_2\text{O}_4^{2-}$ ions
 (C) Mixture contains SO_3^{2-} ions
 (D) Mixture contains both CO_3^{2-} and $\text{C}_2\text{O}_4^{2-}$ ions
- Q.15** An inorganic mixture contains Ag^+ and SO_4^{2-} radicals. The suitable reagent to detect SO_4^{2-} ion is :
- (A) BaCl_2 (B) $\text{Ba}(\text{NO}_3)_2$
 (C) BaSO_4 (D) BaS
- Q.16** Starch + KI paper is turned blue by-
- (A) SO_3^- (B) SO_4^{2-}
 (C) NO_2^- (D) NO_3^-
- Q.17** A deep brown gas is formed by mixing two colourless gases which are :
- (A) NO_2 and O_2 (B) N_2 and NO
 (C) NH_3 and HCl (D) NO and O_2
- Q.18** A salt when heated evolves a coloured gas which bleaches moist litmus paper. The gas is
- (A) Cl_2 (B) Br_2
 (C) NO_2 (D) I_2
- Q.19** Iodide can be obtained from NaI solution by the action of-
- (A) Cl_2 (B) NO_2
 (C) Soluble Cl^- (D) Soluble Br^-
- Q.20** Solution of a compound X reacts with AgNO_3 solution to form a white precipitate of Y which dissolves in NH_4OH to give a complex Z. When Z is treated with dil. HNO_3 , Y reappears. The compound X is-
- (A) NaCl (B) CH_3Cl
 (C) NaBr (D) NaI
- Q.21** During chromyl chloride test, the orange red vapours are of-
- (A) CrOCl_2 (B) ClO_2
 (C) CrO_2Cl_2 (D) Cr_2OCl_2
- Q.22** A white inorganic salt, on heating in a dry test tube turns yellow and on cooling becomes white. The probable salt is :
- (A) ZnO (B) SnO_2
 (C) PbO_2 (D) Sb_2O_3
- Q.23** A salt having BO_3^{-3} on burning with alcohol and conc. H_2SO_4 gives edge flame :
- (A) Green (B) Yellow
 (C) Red (D) White
- Q.24** Aqueous solution of an acetate salt when treated with ferric chloride solution a blood red solution is formed, due to formation of :
- (A) $\text{Fe}(\text{CNS})_3$ (B) $(\text{CH}_3\text{COO})_3\text{Fe}$
 (C) $(\text{CH}_3\text{COO})_2\text{Fe}$ (D) $\text{Fe}(\text{OH})_3$
- MORE THAN ONE OPTION CORRECT :**
- Q.25** Which of the following turn/s acidified $\text{K}_2\text{Cr}_2\text{O}_7$ paper green ?
- (A) SO_2 (B) H_2S
 (C) Br_2 (D) CO_2
- Q.26** Chromyl chloride fails in presence of :
- (A) Ag (B) Sn
 (C) Iodide (D) None of these.
- Q.27** In layer test of I^- ions, the oxidising agent(s) can be :
- (A) Chlorine water (B) Conc. HNO_3
 (C) Sodium hypochlorite (D) Chloramine-T
- Q.28** Heating of oxalic acid with conc. H_2SO_4 gives
- (A) CO (B) CO_2
 (C) SO_2 (D) SO_3
- Q.29** Lime water is turned milky by-
- (A) CO_2 (B) SO_2
 (C) NO_2 (D) SO_3
- Q.30** Which of the following react/s with AgCl ?
- (A) Na_2CO_3 (B) NaNO_3
 (C) NH_4OH (D) $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$



CHEMISTRY IIT JEE (CLASS TEST - 10) (INORGANIC) ANSWER KEY

Name : Roll No. :

	A	B	C	D	E		A	B	C	D	E		A	B	C	D	E
1	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	11	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	21	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
2	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	12	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	22	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
3	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	13	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	23	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
4	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	14	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	24	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
5	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	15	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	25	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
6	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	16	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	26	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
7	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	17	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	27	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
8	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	18	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	28	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
9	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	19	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	29	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
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CHEMISTRY IIT JEE (CLASS TEST - 10) (INORGANIC) ANSWER KEY

Que.	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15
Ans.	D	A	C	B	C	A	B	A	A	B	A	B	B	D	B
Que.	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30
Ans.	C	D	A	A	A	C	A	A	B	A,B	A,B,C	A,B,C	A,B	A,B	A,B

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Que.	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30
Ans.	C	D	A	A	A	C	A	A	B	A,B	A,B,C	A,B,C	A,B	A,B	A,B

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Que.	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30
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Que.	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30
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Que.	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30
Ans.	C	D	A	A	A	C	A	A	B	A,B	A,B,C	A,B,C	A,B	A,B	A,B