

Dear student following is an Easy level [●●○○○] test paper. Any 5 correct attempts in 9 minutes would be a satisfactory performance:

- Q.1 Polarising action of Cd^{2+} on anions is stronger than that of Ca^{2+} because :**
- (A) The charges of the ions are same
 (B) Their radii are same ($\text{Ca}^{2+} = 0.104 \text{ nm}$; $\text{Cd}^{2+} = 0.099 \text{ nm}$)
 (C) The Ca^{2+} ion has a noble-gas electron configuration, and the Cd^{2+} ion, an 18-electron configuration of its outer shell.
 (D) All are correct.
- Q.2 Among the following species, identify the isostructural pairs :**
 NF_3 , NO_3^- , BF_3 , H_3O^+ , HN_3
- (A) [NF_3 , NO_3^-] and [BF_3 , H_3O^+]
 (B) [NF_3 , HN_3] and [NO_3^- , BF_3]
 (C) [NF_3 , H_3O^+] and [NO_3^- , BF_3]
 (D) [NF_3 , H_3O^+] and [HN_3 , BF_3]
- Q.3 Which of the following statements is correct for CsBr_3 ?**
- (A) It is covalent compound
 (B) It contains Cs^{3+} and Br^-
 (C) It contains Cs^+ and Br_3^-
 (D) It contains Cs^+ , Br^- and lattice Br_2 molecule
- Q.4 Which of the following has zero dipole moment ?**
- (A) CO (B) SO_2
 (C) SO_3 (D) H_2O
- Q.5 Which substance has a dipole moment?**
- (A) CCl_4 (B) CH_2Cl_2
 (C) C_2Cl_2 (D) C_2Cl_4
- Q.6 The bond angle in H_2O is 105° and in H_2S it is 90° . It is due to :**
- (A) The larger size of S atom as compared to O atom which minimises repulsion and allows the bonds in H_2S to be purely p-type.
 (B) Liquid state of H_2O as compared to gaseous state of SO_2 .
 (C) Both are correct
 (D) Non is correct
- Q.7 PCl_5 has a shape of trigonal bipyramid whereas IF_5 has the shape of a square pyramid. It is due to :**
- (A) Presence of unshared electron pair on I which is oriented so as to minimise repulsion while P in PCl_5 has no unshared pair.
 (B) Octet of P is complete while that of I is incomplete.
 (C) P and I are of different groups.
 (D) F and Cl have different extent of repulsion.
- Q.8 Which has maximum number of lone pairs of electrons on the central atom ?**
- (A) XeF_2 (B) H_3O^+
 (C) XeF_4 (D) XeF_6
- Q.9 Select correct statement :**
- (A) Cations with 18-electron shell have greater polarising power than the cations with 8-electron shell
 (B) Inner electrons have poor shielding effect on the nucleus and thus electronegativity of the 18-electron shell is increased
 (C) CuCl is covalent and NaCl is ionic
 (D) All are correct statements.

CHEMISTRY IIT JEE (CLASS TEST - 3) (INORGANIC) ANSWER KEY

Name : Roll No. :

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1	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	4	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	7	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
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CHEMISTRY IIT JEE (CLASS TEST - 3) (INORGANIC) ANSWER KEY

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Ans.	C	C	C	C	B	A	A	A	D

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Ans.	C	C	C	C	B	A	A	A	D

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Ans.	C	C	C	C	B	A	A	A	D