

Please read following short write up and answer subsequent questions :

Dimethyl glyoxal is an organic compound which react with hydroxyl amine and form dimethylglyoxime reagent. Nickel sulphate react with dimethyl glyoxime reagent in ammoniacal solution & gives rose red colour. This provide an important test for the qualitative analysis of nickel.

**Q.1 Product form which is rose red in colour colouration is due to -**

- (A) Hydrogen Bonding (B) due to formation of complex  $Ni^{+2}$   
(C) d-d Transition (D) due to formation of complex with  $SO_4^{-2}$

**Q.2 Hybridization sate of nickel in rose red complex & geometry around it is -**

- (A)  $sp^3$ , Tetrahedral (B)  $dsp^2$ , Sq. Planar (C)  $sp^3d^2$ , Octahedral (D)  $d^2sd^3$ , Octahedral  
(E)  $sp^2d$ , Sq. Planar

**Q.3 Spin only magnetic momentum of product is -**

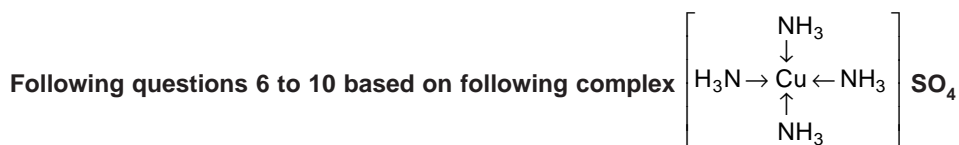
- (A)  $1 \times 9.274 \times 10^{-21}$  erg/gauss $^{-1}$  (B)  $2 \times 9.274 \times 10^{-24}$  erg/gauss $^{-1}$   
(C)  $\frac{1}{2} \times 9.274 \times 10^{-21}$  erg/gauss $^{-1}$  (D) zero

**Q.4 Which type of bond must be present in rose red product -**

- (A) Covalent & Co-ordinate bond (B) Covalent Co-ordinate & inter molecular Hydrogen bond  
(C) Covalent Co-ordinate & intra molecular H bond (D) Inter & Intra molecular Hydrogen bond

**Q.5 Correct IUPAC name of rose red product is -**

- (A) bisdimethylglyoximatnickel (II) sulphate (B) trisdimethylglyoximatnickel (II) sulphate  
(C) bisdimethylglyoximatnickel (II) (D) bisdimethyl glyoximatnickelate (II)



**Q.6 The complex is obtained by.**

- (A) Mixing equimolar amounts of  $\text{CuSO}_4$  and  $(\text{NH}_4)_2 \text{SO}_4$   
(B) Adding concentrated ammonia solution to a solution of  $\text{CuSO}_4$   
(C) Adding ammonia to copper turnings followed by acidification with  $\text{H}_2\text{SO}_4$   
(D) Adding  $(\text{NH}_4)_2 \text{SO}_4$  to a mixture of copper turnings in  $\text{H}_2\text{SO}_4$

**Q.7 The colour of the above complex is.**

- (A) Blue (B) Green (C) Red (D) Violet

**Q.8 The oxidation number of copper in the above complex.**

- (A) 0 (B) 1 (C) 2 (D) 4

**Q.9** Ammonia forms the complex ion  $[\text{Cu}(\text{NH}_3)_4]^{2+}$  with copper ions in alkaline solutions but not in acidic solutions. What is the reason for it ?

- (A) In acidic solutions, protons coordinate with ammonia molecules forming  $\text{NH}_4^+$  ions.  
 (B) In alkaline solutions insoluble  $\text{Cu}(\text{OH})_2$  is precipitated which is soluble in excess of any alkali.  
 (C) Copper hydroxide is an amphoteric substance  
 (D) In acidic solutions hydration protects copper ions.

**Q.10** The effective atomic number of central Cu (At. No. 29) metal in  $[\text{Cu}(\text{NH}_3)_4] \text{SO}_4$  is.

- (A) 29 (B) 30 (C) 35 (D) 36

**Q.No. 11 is based on assertion and reason concept. Select the correct answer :**

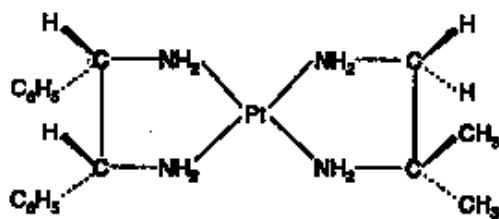
**Q.11 Assertion :** Octahedral complex  $[\text{Co}(\text{edta})]^{-4}$  show optical isomerism.

**Reason :**  $\text{EDTA}^{-4}$  in this complex is hexadentate ligand.

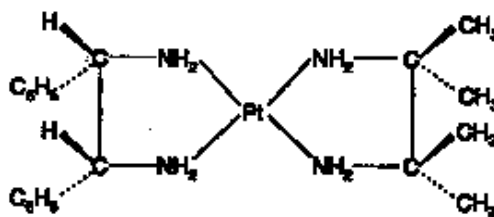
- (A) If both A and R are correct and R is correct explanation of A.  
 (B) If both A and R are correct but R is not the explanation of A.  
 (C) If A is correct but R is wrong.  
 (D) If A is wrong and R is correct.  
 (E) If A and R both are incorrect.

**Following questions have one correct answer. Select the correct answer :**

**Q.12** Given two structure are represent square planar complex of platinum :



(I)



(II)

- (A) I optically active but II is not  
 (B) II optically active but I is not.  
 (C) Both are optically active  
 (D) I & II square planar complex so donot optical isomerism

**Q.13** In Zeigler - Natta polymerisation of ethylene, the active species is.

- (A)  $\text{AlCl}_3$  (B)  $\text{Et}_3\text{Al}$  (C)  $\text{CH}_2\text{CH}_2$  (D)  $\text{Ti}^{\text{III}}$

**Q.14** Which of the following will exhibit optical isomerism.

- (A)  $[\text{Cr}(\text{en})(\text{H}_2\text{O})_4]^{3+}$  (B)  $[\text{Cr}(\text{en})_3]^{3+}$   
 (C)  $\text{trans-}[\text{Cr}(\text{en})(\text{Cl})_2(\text{NH}_3)_2]^+$  (D)  $[\text{Cr}(\text{NH}_3)_6]^{3+}$

**Q.15** The complex ion which has no 'd' electrons in the central metal atom is. (At. No. Cr = 24, Mn = 25, Fe = 26, Co = 27)

- (A)  $[\text{MnO}_4]^-$  (B)  $[\text{Co}(\text{NH}_3)_6]^{3+}$  (C)  $[\text{Fe}(\text{CN})_6]^{3-}$  (D)  $[\text{Cr}(\text{H}_2\text{O})_6]^{3+}$

**Q.16**  $\text{AgCl}$  is soluble in  $\text{NH}_4\text{OH}$  solution. The solubility is due to formation of.

- (A)  $\text{AgOH}$  (B)  $\text{Ag}_2\text{O}$  (C)  $[\text{Ag}(\text{NH}_3)_2]^+$  (D)  $\text{NH}_4\text{Cl}$

**Q.17** Which of the following will give four isomers.

- (A)  $[\text{Co}(\text{en})(\text{NH}_3)_2\text{Cl}_2]\text{Cl}$  (B)  $[\text{Co}(\text{PPh}_3)_2(\text{NH}_3)_2\text{Cl}_2]\text{Cl}$   
 (C)  $[\text{Co}(\text{en})_3]\text{Cl}_3$  (D)  $[\text{Co}(\text{en})_2\text{Cl}_2]\text{Br}$

- Q.18** Which of the following is expected to be a paramagnetic complex  
 (A)  $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$  (B)  $[\text{Ni}(\text{CO})_4]$  (C)  $[\text{Zn}(\text{NH}_3)_4]^{2+}$  (D)  $[\text{Co}(\text{NH}_3)_6]^{3+}$
- Q.19** IUPAC name of  $\text{Na}_3[\text{Co}(\text{NO}_2)_6]$  is.  
 (A) Sodium cobaltinitrite (B) Sodium hexanitrito cobaltate (III)  
 (C) Sodium hexanitro cobalt (III) (D) Sodium hexanitro cobaltate (III)
- Q.20** The oxidation state of Fe in brown ring complex  $[\text{Fe}(\text{H}_2\text{O})_5\text{NO}]\text{SO}_4$  is.  
 (A) + 1 (B) + 2 (C) + 3 (D) + 4
- Q.21** The oxidation state and coordination number of Cr in  $[\text{Cr}(\text{C}_2\text{O}_4)_3]^{3-}$  ion are.  
 (A) + 6 and 3 (B) + 3 and 6 (C) + 6 and 6 (D) + 3 and 3
- Q.22** Hypo is used in photography because it is.  
 (A) Strong oxidising agent (B) Strong reducing agent  
 (C) Strong complexing agent (D) None of the above
- Q.23** The characteristic oxidation number of atoms in free metals is.  
 (A) Minus one (B) One (C) Any number (D) Zero
- Q.24** Ligands in a complex salt are.  
 (A) Anions linked by coordinate bonds to a central metal ion  
 (B) Cations linked by coordinate bonds to a central metal ion  
 (C) Molecules linked by coordinate bonds to a central metal ion  
 (D) Ions or molecules linked by coordinate bonds to a central metal ion
- Q.25** Amongst  $\text{Ni}(\text{CO})_4$ ,  $[\text{Ni}(\text{CN})_4]^{2-}$  and  $\text{NiCl}_4^{2-}$ .  
 (A)  $\text{Ni}(\text{CO})_4$  and  $\text{NiCl}_4^{2-}$  are diamagnetic and  $[\text{Ni}(\text{CN})_4]^{2-}$  is paramagnetic  
 (B)  $\text{NiCl}_4^{2-}$  and  $[\text{Ni}(\text{CN})_4]^{2-}$  are diamagnetic and  $\text{Ni}(\text{CO})_4$  is paramagnetic  
 (C)  $\text{Ni}(\text{CO})_4$  and  $[\text{Ni}(\text{CN})_4]^{2-}$  are diamagnetic and  $\text{NiCl}_4^{2-}$  is paramagnetic  
 (D)  $\text{Ni}(\text{CO})_4$  is diamagnetic and  $\text{NiCl}_4^{2-}$  and  $[\text{Ni}(\text{CN})_4]^{2-}$  are paramagnetic.
- Q.26**  $\text{K}_4[\text{Fe}(\text{CN})_6]$  is called.  
 (A) Potassium hexacyanoferrate (II) (B) Potassium ferricyanide  
 (C) Potassium hexacyanoferrate (III) (D) Prussian blue
- Q.27** Wilkinson's catalyst,  $(\text{Ph}_3\text{P})_3\text{RhCl}$  is used for.  
 (A) Hydrogenation of oils (B) Hydrogenation of alkynes  
 (C) Hydrogenation of alkenes (D) Polymerisation of alkenes
- Q.28** cis - Platin used for treatment of cancer is.  
 (A)  $[\text{Pt}(\text{NH}_3)_4\text{Cl}_2]$  (B)  $[\text{Pt}(\text{NH}_3)_2\text{Cl}_2]$  (C)  $[\text{Pt}(\text{NH}_3)_2\text{Cl}_4]$  (D)  $[\text{Pt}(\text{NH}_3)_3\text{Cl}_3]$
- Q.29** EDTA is used for the estimation of.  
 (A)  $\text{Na}^+$  and  $\text{K}^+$  ions (B)  $\text{Cl}^-$  and  $\text{Br}^-$  ions (C)  $\text{Cu}^{2+}$  and  $\text{Ag}^+$  ions (D)  $\text{Ca}^{2+}$  and  $\text{Mg}^{2+}$  ions

- Q.30** The stability constants of the complexes formed by a metal ion ( $M^{2+}$ ) with  $NH_3$ ,  $CN^-$ ,  $H_2O$  and en are of the order  $10^{11}$ ,  $10^{27}$ ,  $10^{15}$ , and  $10^8$  respectively. Then.
- (A) en is the strongest ligand  
 (B)  $CN^-$  is the strongest ligand  
 (C) These values cannot predict the strength of the ligand  
 (D) All ligands are equally strong
- Q.31** The complex ions  $[Fe(CN)_6]^{3-}$  and  $[Fe(CN)_6]^{4-}$ .
- (A) Are both octahedral and paramagnetic  
 (B) Are both octahedral and diamagnetic  
 (C) Have same structure but opposite magnetic character  
 (D) Have different structures but same magnetic character
- Q.32** The shape of the complex  $[Ni(CO)_4]$  is.
- (A) Square planar      (B) Tetrahedral      (C) Octahedral      (D) None of these
- Q.33** Which of the following is a complex of a metal other than transition metal.
- (A) Haemoglobin      (B) Chlorophyll      (C) Ferrocene      (D) Vitamin  $B_{12}$
- Q.34** The correct name for the complex  $[Cr(NH_3)_6][Co(C_2O_4)_3]$  is.
- (A) Hexaamminechromium (III) trioxalatocobalt (III)  
 (B) Hexaamminechromate (III) trioxalatocobaltate (III)  
 (C) Hexaamminechromium (III) trioxalatocobaltate (III)  
 (D) Hexaamminechromate (III) trioxalatocobalt (III)
- Q.35** Which one of the following is not an organometallic compound.
- (A)  $(C_2H_5)_2Zn$       (B)  $CH_3Al(OCH_3)_2$       (C)  $B(OCH_3)_3$       (D)  $Ni(CO)_4$
- Q.36** In the compound lithiumtetrahydroaluminate, the ligand is.
- (A)  $H^+$       (B)  $H$       (C)  $H^-$       (D) None of these
- Q.37** A complex compound of cobalt has molecular formula containing five  $NH_3$  molecules, one nitro group and two chlorine atoms for one Co atom. One mole of this compound produces three mole ions in aq. solution. On reacting with excess of  $AgNO_3$  solution, two moles of  $AgCl$  get precipitated. The ionic formula of the compound is.
- (A)  $[Co(NH_3)_4NO_2Cl] [(NH_3)Cl]$       (B)  $[Co(NH_3)_5Cl] [Cl(NO_2)]$   
 (C)  $[Co(NH_3)_5(NO_2)] Cl_2$       (D)  $[Co(NH_3)_5] [(NO_2)Cl_2]$
- Q.38** A solution containing 2.674 gm of  $CoCl_3 \cdot 6NH_3$  was passed through a cation exchanger. The solution obtained gave 4.305 gm of  $AgCl$  on treatment with  $AgNO_3$ . The formula of the complex is
- (A)  $[Co(NH_3)_5Cl]Cl_2NH_3$       (B)  $[CoCl_3](NH_3)_6$       (C) Both (A) and (B)      (D)  $[Co(NH_3)_6]Cl_3$
- Q.39** Stability constant is more for the complex.
- (A)  $[Ag(NH_3)_2]^+$       (B)  $[Cu(NH_3)_4]^{2+}$       (C)  $[Ag(CN)_2]^{-1}$       (D)  $[Co(NH_3)_6]^{3+}$
- Q.40** From the stability constants (hypothetical values), given below, predict which one is the stable complex.
- (i)  $Cu^{2+} + 4NH_3 \rightleftharpoons [Cu(NH_3)_4]^{2+}$ ,  $K = 4.5 \times 10^{11}$       (ii)  $Cu^{2+} + 4CN^- \rightleftharpoons [Cu(CN)_4]^{2-}$ ,  $K = 2.0 \times 10^{27}$   
 (iii)  $Cu^{2+} + 2en \rightleftharpoons [Cu(en)_2]^{2+}$ ,  $K = 3.0 \times 10^{15}$       (iv)  $Cu^{2+} + 4H_2O \rightleftharpoons [Cu(H_2O)_4]^{2+}$ ,  $K = 9.5 \times 10^8$
- (A)  $[Cu(NH_3)_4]^{2+}$       (B)  $[Cu(CN)_4]^{2-}$       (C)  $[Cu(en)_2]^{2+}$       (D)  $[Cu(H_2O)_4]^{2+}$

**Q.41 Match List 1 and List 2, (the basis of matching being complex formation in qualitative or quantitative analysis of ions) and select the correct answer using the codes given below the lists**

List 1					List 2				
Ions involved					Complexing agent				
1. Ni <sup>2+</sup>					A. EDTA				
2. Zn <sup>2+</sup>					B. Sodium nitroprusside				
3. Cu <sup>2+</sup>					C. Ammonia				
4. S <sup>2-</sup>					D. Dimethylglyoxime				
	1	2	3	4		1	2	3	4
(A)	A	B	D	C	(B)	D	A	C	B
(C)	C	D	B	A	(D)	D	C	A	B

**Q.42 Match List I with List II and select the correct answer using the codes given below the lists.**

List I					List II						
Complex Ions					Magnetic Moment ( $\mu$ ) Bohr Magnetons						
1. [Fe(CN) <sub>6</sub> ] <sup>4-</sup>					A. 1.73						
2. [Ti(H <sub>2</sub> O) <sub>6</sub> ] <sup>3+</sup>					B. 5.93						
3. [Cr(H <sub>2</sub> O) <sub>6</sub> ] <sup>3+</sup>					C. 0.00						
4. [Ni(H <sub>2</sub> O) <sub>6</sub> ] <sup>2+</sup>					D. 2.83						
5. [FeF <sub>6</sub> ] <sup>3-</sup>					E. 3.88						
	1	2	3	4	5		1	2	3	4	5
(A)	A	B	C	D	E	(B)	C	A	E	D	B
(C)	B	C	D	E	A	(D)	D	E	A	B	C

**Q.43 In nitroprusside ion, the iron and NO exist as Fe(II) and NO<sup>+</sup> rather than Fe(III) and NO. These forms can be differentiated by-**

- (A) Estimating the concentration of ion  
(B) Measuring the concentration of CN<sup>-</sup>  
(C) Measuring the solid state magnetic moment  
(D) Thermally decomposing the compound

**Q.44 Which order is correct in spectrochemical series of ligands-**

- (A) Cl<sup>-</sup> > F<sup>-</sup> < C<sub>2</sub>O<sub>4</sub><sup>2-</sup> < NO<sub>2</sub><sup>-</sup> < CN<sup>-</sup>  
(B) CN<sup>-</sup> < C<sub>2</sub>O<sub>4</sub><sup>2-</sup> < Cl<sup>-</sup> > NO<sub>2</sub><sup>-</sup> < F<sup>-</sup>  
(C) C<sub>2</sub>O<sub>4</sub><sup>2-</sup> < F<sup>-</sup> < Cl<sup>-</sup> > NO<sub>2</sub><sup>-</sup> < CN<sup>-</sup>  
(D) F<sup>-</sup> < Cl<sup>-</sup> < NO<sub>2</sub><sup>-</sup> < CN<sup>-</sup> < C<sub>2</sub>O<sub>4</sub><sup>2-</sup>

**Q.45 K<sub>3</sub>CoF<sub>6</sub> is high spin complex. What is the hybrid state of Co atom in this complex-**

- (A) sp<sup>3</sup>d  
(B) sp<sup>3</sup>d<sup>2</sup>  
(C) d<sup>2</sup>sp<sup>3</sup>  
(D) dsp<sup>2</sup>

**Q.46 A complex with the composition [MA<sub>2</sub>B<sub>2</sub>] X<sub>2</sub> is found to have no geometrical isomers. Both A and B are monodentate ligands. The structure of the complex is-**

- (A) Linear  
(B) Square planar  
(C) Tetrahedral  
(D) Octahedral

**Q.47 The probable formula for Prussian blue is-**

- (A) Fe<sub>3</sub>[Fe(CN)<sub>6</sub>]<sub>2</sub>  
(B) Fe<sub>2</sub>[Fe(CN)<sub>6</sub>]<sub>3</sub>  
(C) Fe<sub>4</sub>[Fe(CN)<sub>6</sub>]<sub>3</sub>  
(D) Fe<sub>3</sub>[Fe(CN)<sub>6</sub>]<sub>4</sub>

**Q.48 Addition of KI to Hgl<sub>2</sub> forms complex K<sub>2</sub>Hgl<sub>4</sub> having ...**

- (A) Red colour  
(B) Blue colour  
(C) Violet colour  
(D) Colourless nature

**Q.49 Which of the following is  $\pi$ -acid ligand-**

- (A) NH<sub>3</sub>  
(B) CO  
(C) F<sup>-</sup>  
(D) Ethylene diamine

**Q.50 Organometallic compound used in the purification of its metals is-**

- (A) Ni(CO)<sub>4</sub>  
(B) Pb(C<sub>2</sub>H<sub>5</sub>)<sub>4</sub>  
(C) Li - C<sub>4</sub>H<sub>9</sub>  
(D) Na<sub>2</sub>[Ni(CN)<sub>4</sub>]

- Q.51 The value of stability constant depends upon-**  
 (A) The charge on the central metal ion (B) Nature of the ligand  
 (C) Chelation (D) All of the above
- Q.52 Ferrocene is-**  
 (A)  $\text{Fe}(\eta^5 - \text{C}_5\text{H}_5)_2$  (B)  $\text{Cr}(\eta^6 - \text{C}_6\text{H}_6)_2$  (C)  $\text{K}[\text{PtCl}_3 (\eta^2 - \text{C}_2\text{H}_4)]$  (D)  $(\text{C}_2\text{H}_5)_4\text{Pb}$
- Q.53 The complex formed during the fixing of negative in photography is-**  
 (A)  $[\text{Ag}(\text{S}_2\text{O}_3)_2]^{-1}$  (B)  $[\text{Ag}(\text{S}_2\text{O}_3)_2]^{2-}$  (C)  $[\text{Ag}(\text{S}_2\text{O}_3)_2]^{3-}$  (D) None of these
- Q.54 Sodium nitroprusside is-**  
 (A)  $\text{Na}[\text{Fe}(\text{CN})_5\text{NO}]$  (B)  $\text{Na}_3[\text{Fe}(\text{CN})_5\text{NO}]$  (C)  $\text{Na}[\text{Fe}(\text{CN})_4(\text{NO})_2]$  (D)  $\text{Na}_2[\text{Fe}(\text{CN})_5\text{NO}]$
- Q.55 The number of moles of  $\text{BaSO}_4$  precipitated per mole of Schewtizer's reagent on adding excess of  $\text{BaCl}_2$  are-**  
 (A) 2 (B) 0 (C) 1 (D) None of these
- Q.56 Ammonia will not form complex ions with-**  
 (A)  $\text{Ag}^+$  (B)  $\text{Cd}^{2+}$  (C)  $\text{Cu}^{2+}$  (D)  $\text{Pb}^{2+}$ .
- Q.57 A similarity between optical and geometrical isomerism is that -**  
 (A) Each gives equal number of isomers for a given compound  
 (B) If in a compound one is present then so is that other  
 (C) Both are included in stereoisomerism  
 (D) They have no similarity
- Q.58 A square planar complex is formed by hybridization of which atomic orbitals ?**  
 (A) s,  $p_x$ ,  $p_y$ ,  $d_{yz}$  (B) s,  $p_x$ ,  $p_y$ ,  $d_{x^2-y^2}$  (C) s,  $p_x$ ,  $p_y$ ,  $d_{z^2}$  (D) s,  $p_x$ ,  $p_y$ ,  $d_{xy}$
- Q.59 Fac-Mer isomerism is associated with which of the following general formula -**  
 (A)  $\text{M}(\text{AA}')_2$  (B)  $\text{M}(\text{AA})_3$  (C)  $\text{MABCD}$  (D)  $\text{MA}_3\text{B}_3$
- Q.60 The complex compounds formed when KCN solution is added to a solution containing both  $\text{Cu}^{2+}$  and  $\text{Cd}^{2+}$  ions are**  
 (A)  $\text{K}_2[\text{Cu}(\text{CN})_4]$ ,  $\text{K}_2[\text{Cd}(\text{CN})_4]$  (B)  $\text{K}_3[\text{Cu}(\text{CN})_4]$ ,  $\text{K}_2[\text{Cd}(\text{CN})_4]$   
 (C)  $\text{K}_3[\text{Cu}(\text{CN})_4]$ ,  $\text{K}_3[\text{Cd}(\text{CN})_4]$  (D)  $\text{K}_2[\text{Cu}(\text{CN})_4]$ ,  $\text{K}_3[\text{Cd}(\text{CN})_4]$



## IIT JEE CLASS TEST - 5 (INORGANIC) ANSWER KEY

Name : ..... Roll No. : .....

Father's Name : ..... Phone No : ..... Batch : .....

Address : ..... Date : .....

1	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	21	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	41	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
2	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	22	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	42	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
3	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	23	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	43	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
4	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	24	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	44	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
5	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	25	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	45	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
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	No. of Correct Ans.	Marks
Total		

*Total  
Marks*

**Rank**

