

Dear student following is Moderate level [●●●○○] test paper. Any 6 correct attempts in 10 minutes would be a satisfactory performance:

- Q.1 The standard electrode potential of transition element is-**
 (A) Sufficiently negative
 (B) Positive
 (C) Zero
 (D) None of these
- Q.2 d-block elements generally form-**
 (A) Covalent hydrides
 (B) Metallic hydrides
 (C) Interstitial hydrides
 (D) Salt like hydrides
- Q.3 Hardness of transition elements in comparison to alkali metals is due to-**
 (A) Large atomic size
 (B) Metallic bonding
 (C) Covalent bonding
 (D) High ionisation energy
- Q.4 Which one of the following electronic configurations is that of transition elements ?**
 (A) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^5$
 (B) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^1$
 (C) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^2 4s^2$
 (D) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$
- Q.5 Which one of the following statements is incorrect about transition elements ?**
 (A) The last electron enters the d-orbitals
 (B) Their properties lie in between those of s-block elements and p-block elements
 (C) The transition element with smallest atomic number is scandium
 (D) Lanthanum is not a transition element but is a lanthanide
- Q.6 The calculated value of magnetic moment of $_{23}V^{3+}$ is-**
 (A) $1.73 \mu_B$ (B) $2.83 \mu_B$
 (C) $3.87 \mu_B$ (D) $4.9 \mu_B$
- Q.7 Compared to Cu^{2+} having $3d^9$ configuration, Cu^+ having $3d^{10}$ configuration**
 (A) Is more stable
 (B) Is less stable
 (C) Is equally stable
 (D) Stability depends upon the nature of copper salt.
- Q.8 In the dichromate dianion :**
 (A) 4 Cr - O bonds are equivalent
 (B) 6 Cr - O bonds are equivalent
 (C) All Cr - O bonds are equivalent
 (D) All Cr - O bonds are non-equivalent
- Q.9 The electronic configuration of an element is $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^1$. This represents its-**
 (A) Excited state (B) Ground state
 (C) Cationic form (D) Anionic form
- Q.10 Standard electrode potential data are useful for understanding the suitability of an oxidant in a redox titration. Some half cell reactions and their standard potentials are given below**
 $MnO_4^-(aq.) + 8H^+(aq.) + 5e^- \rightarrow Mn^{2+}(aq.) + 4H_2O(l)$
 $E^\circ = 1.51 V$
 $Cr_2O_7^{2-}(aq.) + 14H^+(aq.) + 6e^- \rightarrow 2Cr^{3+}(aq.) + 7H_2O(l)$
 $E^\circ = 1.38 V$
 $Fe^{3+}(aq.) + e^- \rightarrow Fe^{2+}(aq.)$ $E^\circ = 0.77V$
 $Cl_2(g) + 2e^- \rightarrow 2Cl^-(aq.)$ $E^\circ = 1.40V$
Identify the only incorrect statement regarding the quantitative estimation of aqueous $Fe(NO_3)_2$
 (A) MnO_4^- can be used in aqueous HCl
 (B) $Cr_2O_7^{2-}$ can be used in aqueous HCl
 (C) MnO_4^- can be used in aqueous H_2SO_4
 (D) $Cr_2O_7^{2-}$ can be used in aqueous H_2SO_4

CHEMISTRY IIT JEE (CLASS TEST - 6) (INORGANIC) ANSWER KEY

Name : Roll No. :

	A	B	C	D	E		A	B	C	D	E		A	B	C	D	E
1	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	4	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	7	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
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CHEMISTRY IIT JEE (CLASS TEST - 6) (INORGANIC) ANSWER KEY

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Ans.	A	C	C	C	D	B	B	B	B	A

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